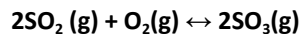


Acids Unit: Redox Application

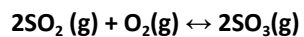
1. The following reaction takes place in a closed container.



The oxidizing agent in the reaction is _____

a	$\text{SO}_2(\text{g})$
b	$\text{O}_2(\text{g})$
c	$\text{SO}_3(\text{g})$
d	This is not an oxidation reduction reaction

2. The following reaction takes place in a closed container.



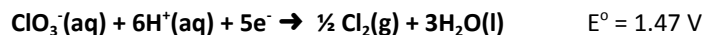
The reducing agent in the reaction is _____

a	$\text{SO}_2(\text{g})$
b	$\text{O}_2(\text{g})$
c	$\text{SO}_3(\text{g})$
d	This is not an oxidation reduction reaction

3. Hydrogen gas is produced when magnesium metal reacts with 1.0 mol/L of _____.

a	$\text{CH}_3\text{OH}(\text{aq})$
b	$\text{C}_2\text{H}_5\text{OH}(\text{aq})$
c	$\text{CH}_3\text{COOH}(\text{aq})$
d	$\text{C}_6\text{H}_{12}\text{O}_6(\text{aq})$

4. Consider the following half reactions



A student mixes acidified solutions of $\text{KCl}(\text{aq})$ and $\text{KClO}_3(\text{aq})$. The FALSE observation is _____

a	Chlorine gas is produced
b	The pH of the mixture decreases.
c	A spontaneous reaction occurs.
d	The concentrations of $\text{ClO}_3^-(\text{aq})$ and $2\text{Cl}^-(\text{aq})$ both decrease.

Solutions:

1. B
2. A
3. C
4. B