

Anode & Cathode

1. During the electroplating of copper spoons, copper is deposited by

a	oxidizing copper ions at the cathode
b	reducing copper ions at the cathode
c	oxidizing copper ions at the anode
d	reducing copper ions at the anode

2. Use this information to answer the question that follows.



This equation represents the reaction in an electrochemical cell.

The half-cell reaction that occurs at the anode is _____

a	$\text{Zn}(\text{s}) \rightarrow 2\text{e}^{-} + \text{Zn}^{2+}(\text{aq})$
b	$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Zn}(\text{s})$
c	$\text{Al}(\text{s}) \rightarrow 3\text{e}^{-} + \text{Al}^{3+}(\text{aq})$
d	$\text{Al}^{3+}(\text{aq}) + 3\text{e}^{-} \rightarrow \text{Al}(\text{s})$

3. In an electrochemical cell, the cathode is the electrode at which electrons _____ the half cell and _____ takes place.

a	Leave	Oxidation
b	Leave	Reduction
c	Enter	Oxidation
d	Enter	Reduction

Solutions:

1. B
2. C
3. D